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[Fluid Phase Equilibria - Journal - Elsevier](#)

Special issue of Fluid Phase Equilibria in memory of Dr. Costa Tsonopoulos. Edited by Clare McCabe, Ioannis Economou, Georgios Kontogeorgis. 7 January 2021. Carbon Capture and Storage . Edited by Professor Clare McCabe. 27 August 2020. Selected papers from the X Brazilian Conference on Applied Thermodynamics (CBTermo 2019) Edited by Eduardo Rocha de Almeida Lima, Márcio Luis Lyra Paredes ...

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Experimental data of vapor – liquid equilibria (VLE) are reported for the binary systems: ethanol (1) + water (2) at 100, 1500, and 2000 kPa. The data were verified for thermodynamic consistency by using different low- and high-pressure tests. The following minimum azeotropes were found for the binary systems: $x_{1,az} = y_{1,az} = 0.899$ and $T = 351.39$ K at 100 kPa, $x_{1,az} = y_{1,az} = 0.857$ and $T = 440$...

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Mats Hillert: Phase equilibria, phase diagrams and phase transformations: their thermodynamic basis. 2. Auflage. Cambridge University Press, Cambridge, UK 2008, ISBN 978-1-139-12921-3. Bruno Predel, Michael Hoch, Monte Pool: Phase Diagrams and Heterogeneous Equilibria: a Practical Introduction. Springer-Verlag, Berlin/Heidelberg 2004, ISBN 978-3-662-09276-7, doi: 10.1007/978-3-662-09276-7

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Gas-phase equilibria. For equilibria in a gas phase, fugacity, f , is used in place of activity. However, fugacity has the dimension of pressure, so it must be divided by a standard pressure, usually 1 bar, in order to produce a dimensionless quantity, f / p^o . An equilibrium constant is expressed in terms of the dimensionless quantity. For ...

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The non-random two-liquid model (abbreviated NRTL model) is an activity coefficient model that correlates the activity coefficients of a compound with its mole fractions in the liquid phase concerned. It is frequently applied in the field of chemical engineering to calculate phase equilibria. The concept of NRTL is based on the hypothesis of Wilson that the local concentration around a ...

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Temkin equation is excellent for predicting the gas phase equilibrium (when organization in a tightly packed structure with identical orientation is not necessary), conversely complex adsorption systems including the liquid-phase adsorption isotherms are usually not appropriate to be represented . 2.1.5. Flory – Huggins isotherm model

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